

A Catalyst and the Rate of Reaction - Activity Sheet

Objective: _____

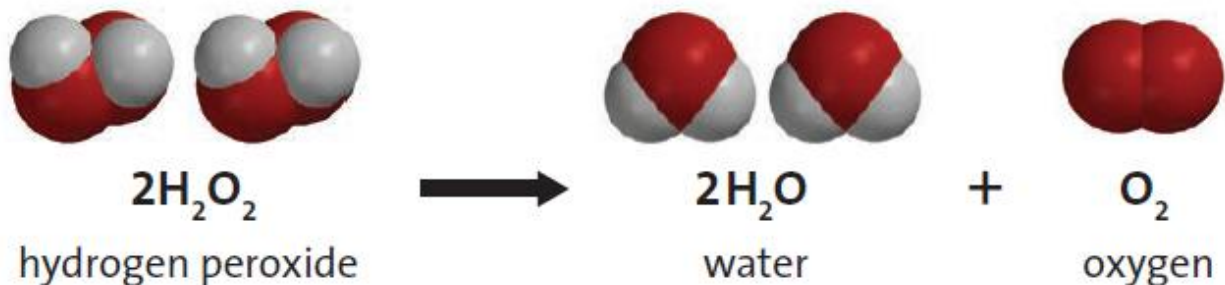
DEMONSTRATION

1. Your teacher showed you videos of two chemistry demonstrations: Elephant's toothpaste and genie in a bottle. Are both of these chemical changes?

How do you know?

EXPLAIN IT WITH ATOMS & MOLECULES

2. Even though the two demonstrations seemed different, the chemical reaction behind both is exactly the same - the decomposition of hydrogen peroxide. Refer to the following equation as you answer the questions below.



Which new substances are created when hydrogen peroxide decomposes?

Each demonstration used a substance called a catalyst. What does a catalyst do in a chemical reaction?

If the catalyst is involved in the chemical reaction, why isn't it included as a product in the chemical equation?

ACTIVITY

Question to Investigate

Can another substance catalyze the decomposition of hydrogen peroxide?

Materials for Each Group

- Graduated cylinder
- Hydrogen peroxide (3%)
- Yeast
- Popsicle stick
- Detergent solution
- Dropper



Procedure

1. Add 10 mL of hydrogen peroxide to a graduated cylinder. Add 1 drop of detergent solution. Swirl gently and watch the solution for any bubbling.
 2. Use the end of a Popsicle stick to add a small amount of yeast to the hydrogen peroxide in the graduated cylinder and swirl.
 3. Place the graduated cylinder on the table and watch for any bubbling.
 4. Hold the graduated cylinder to see if there seems to be any change in temperature.
3. What clues did you have that a chemical reaction occurred in this activity?
4. What is the catalyst in this activity?
5. What evidence do you have that hydrogen peroxide decomposed faster when you added yeast?
6. When writing the chemical equation for this reaction, should yeast be included on the product side of the chemical equation?

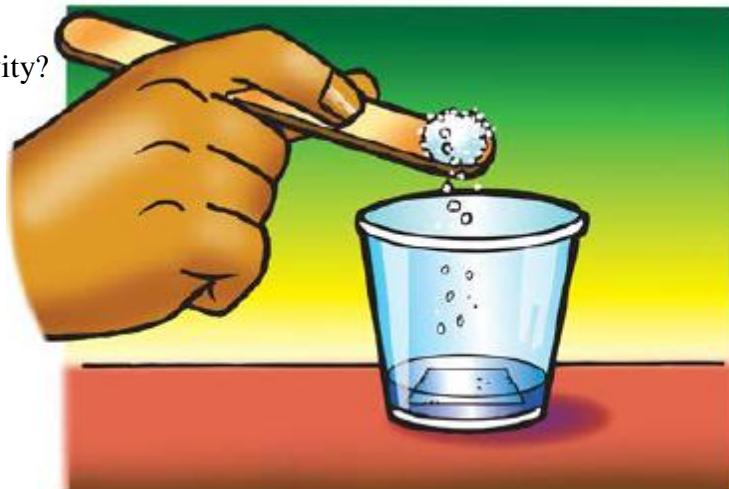
TAKE IT FURTHER

Question to Investigate

What is the catalyst in the following activity?

Materials for Each Group

- Copper II sulfate solution
- Beaker
- Salt
- Piece of aluminum foil
- Thermometer
- Popsicle Stick



Procedure

1. Place the piece of aluminum foil in an empty beaker. Use your fingers or a popsicle stick to push the foil down so that it lays flat and covers the bottom of the beaker.
 2. Add enough copper II sulfate solution to the beaker to cover the aluminum foil in the bottom of the beaker.
 3. Gently swirl the solution for a few seconds and let it stand still. Watch the aluminum for any bubbling or color change.
 4. Use your popsicle stick to place a small amount of salt in the copper II sulfate solution. Gently swirl the solution for a few seconds and let it stand still. Watch for any bubbling or color change.
 5. Carefully place a thermometer in the cup and see if the temperature changes.
7. How do you know that a chemical reaction occurs when a piece of aluminum foil and sodium chloride is placed in copper II sulfate solution?
8. What is the catalyst in this activity?
9. How is adding the salt to the aluminum similar to adding yeast to the hydrogen peroxide?