

## Temperature and the Rate of a Chemical Reaction - Notes Sheet

### Key Concepts:

- ❖ Reactants must be moving fast enough and hit each other hard enough for a \_\_\_\_\_ to take place.
- ❖ Increasing the \_\_\_\_\_ increases the average speed of the \_\_\_\_\_ molecules.
- ❖ As more molecules move faster, the number of molecules moving fast enough to react \_\_\_\_\_, which results in faster \_\_\_\_\_ of \_\_\_\_\_.

### Temperature and the rate of a Chemical Reaction Processing

#### EXPLAIN IT WITH ATOMS & MOLECULES

Does the temperature of the reactants affect the rate of the chemical reaction?

How do you know?

On the molecular level, why do you think the warm solutions react faster than the cold solutions?

Add a picture to demonstrate what happened in your activity: